

LEYKIS, W. E.

LEYKIS D. I.
LEKOVICH, P. D.

PHASE I BOOK EXPLOITATION SOV/2216

5(4)

Sveshchaniye po elektrolizu. 4th, Moscow, 1956.
Trudy... [bornik] (Transactions of the Fourth Conference on Electrochemistry; Collection of Articles) Moscow, Izd-vo AN SSSR, 1959. 868 p. Kzrta slyp... 2,500 copies printed.
Sponsoring Agency: Akademiya nauk SSSR. Otdeleniye khimicheskikh nauk.

Editorial Board: A. M. Fumkin (Resp. Ed.) Academician, O. A. Yasin, Professor, I. I. Zhdanov (Resp. Secretary), B. M. Kabanov, Professor, S. I. Zhdanov (Resp. Secretary), B. M. Kabanov, Professor, I. I. Zhdanov, Doctor of Chemical Sciences, V. V. Loser, P. D. Lekevich, Professor, Z. A. Solov'yeva, V. V. Loser, Professor, and O. M. Florianovich; Ed. of Publishing House M. G. Yegorov; Tech. Ed.: T. A. Prusikova.

PURPOSE: This book is intended for chemical and electrical engineers, physicists, metallurgists and researchers interested in various aspects of electrochemistry.

COVERAGE: The book contains 127 of the 138 reports presented at the Fourth Conference on Electrochemistry sponsored by the Department of Chemical Sciences and the Institute of Physical Chemistry, Academy of Sciences, USSR. The collection contains theories and branches of electrochemical kinetics, double layer, industrial electrolysis, processes in metal electrodes, the end of each electrolytic process, and galvanic cells. The majority of reports not included here have been published in periodical literature. Personalities are mentioned. References are given at the end of most of the articles.

A. A. Zhdanov-Gorkiy-Polytechnic Institute Iseni A. A. Zhdanov. Influence of Aging Processes on the Work of Alkaline-Zinc Elements 768

Lukovits, I. D. Theory of Processes Occurring at Oxide Electrodes of Chemical Sources of Current 773

Rosentavere, S. A., and V. Y. Lavina. Mechanism of the Activation of an Iron Electrode With Small Additions of Nickel Oxides 781

Balshova, M. A., V. A. Ivanov, and L. D. Kovba. Institute of Electrochemistry, Academy of Sciences, USSR, Being Tested Atoms to Study Processes in Chemical Sources of Current 788

Danilov-Bek, V. S., M. Z. Muzes, V. V. Syagruva, and M. V. Yel'manova (Nauchnoissledovatel'skiy institut gorodskoy i sel'skoy avyazki khimicheskoy avyazki SSSR - Scientific Research Institute of Rural and Urban Communications, Pul Ministry of Communications, USSR). Investigation of Fuel

Card 31/34

Shumakovskaya, M. A., and A. Kh. Burahteyev. Institute for Electrochemistry, AS USSR, Moscow. Iron-Carbon Element 801

Levits, D. I. (Institute of Electrochemistry, Academy of Sciences, USSR). Effect of Salt or Oxide Layers Formed in Discharge or Charging Processes on the Passivation of Battery Electrodes 807

Selishkova, S. P., and L. A. Leont'yeva. Influence of Cathodic Polarization at Low Temperatures on the Anode Potential of an Iron Electrode in an Alkaline Solution 811

Discussion [S. A. Gerasim, M. S. Lidorenko, P. P. Yuppets, A. P. Ksenofontov and contributing authors] 814

PART X. ELECTROLYSIS IN THE CHEMICAL INDUSTRY 821

Card 32/34

5(4)

AUTHORS:

SOV/20-124-4-35/67
Vidovich, G. L., Leykis, D. I., Kabanov, B. N.

TITLE:

The Anodic Passivation of Silver in Alkali Solutions
(Anodnaya passivatsiya serebra v rastvorakh shchelochi)

PERIODICAL:

Doklady Akademii nauk SSSR, 1959, Vol 124, Nr 4, pp 855-857
(USSR)

ABSTRACT:

The present paper investigates the passivation of silver in solutions of KOH in the concentration interval of from 0.01 n to 10 n according to the method of measuring the capacity and the ohmic components of the impedance of the electrode at constant current density or constant potential. Impedance was measured by employing the compensation method. If a low positive potential (0.3 v against the oxide-mercury electrode in the same solution) is applied, the dissolution rate of the silver decreases rapidly, and, at the same time, also the capacity of the double layer of the electrode is reduced. These considerable modifications of the intensity of the dissolving current and of the electrode capacity occur during the passage of a smaller quantity of electricity (~ 20 millicoulomb cm^2), which corresponds on the average to about 10 molecular layers of Ag_2O . The capacity decrease

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The Anodic Passivation of Silver in Alkali Solutions

is apparently connected with the semiconductor properties of the Ag_2O . Here it is assumed that the Ag_2O -layer formed covers the electrode in a relatively uniform manner. The data discussed indicate the following: At the very beginning, before a compact Ag_2O -layer is formed on the surface, a solution with primary formation of an oversaturated solution and with a following deposition at the centers of crystallization is formed. After the formation of a compact Ag_2O -layer on the surface of the silver, the $\text{Ag} \rightarrow \text{Ag}_2\text{O}$ process apparently develops on the boundary between two solid phases. During the formation of the compact Ag_2O -layer the rate at which the process $\text{Ag} \rightarrow \text{Ag}_2\text{O}$ develops decreases considerably. This is described by the authors as the first stage of passivation. However, the process develops only during a limited period at this reduced rate. In the case of constant amperage the potential increases considerably after a certain interval of time, and the second stage of passivation begins. At the same time, slowing down of the electrochemical process $\text{Ag} \rightarrow \text{Ag}_2\text{O}$ as a whole increases rapidly. At this stage this is probably the electrochemical adsorption of oxygen on the surface of

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The Anodic Passivation of Silver in Alkali Solutions.

the silver, similar as is the case in the solution and passivation of iron in a lye. These as well as other results discussed here lead to the following conclusion: The anodic passivation of silver in lye takes place in three successive stages, and in each of these stages the rate of oxidation decreases rapidly. The investigation of the mechanism of these stages of passivation is being continued. There are 2 figures and 5 references, 4 of which are Soviet.

ASSOCIATION: Institut elektrokhemii Akademii nauk SSSR
(Institute for Electrochemistry of the Academy of Sciences, USSR)

PRESENTED: October 1, 1958, by A. N. Frumkin, Academician

SUBMITTED: September 11, 1958

Card 3/3

S/076/60/034/008/035/039/XX
B015/B063

AUTHORS: Leykis, D. I., Knots, L. L.
TITLE: Detection of Concentration Polarization by Measuring the
Electrode Impedance
PERIODICAL: Zhurnal fizicheskoy khimii, 1960, Vol. 34, No. 8,
pp. 1872-1874

TEXT: For kinetic studies of electrodic processes it is important to know whether a concentration polarization takes place at the electrode. This problem is usually solved by measuring the component of the electrode impedance in alternating current at different frequencies, since the concentration polarization at the electrode is equivalent to the corresponding values of capacitance and resistance which are connected in parallel or series. The phase shift is 45° , and the impedance component is a linear function of $1/\sqrt{\omega}$ (ω = angular frequency of alternating current). If the capacitance of the double layer (electrode - electrolyte) or the concentration polarization is high, this function becomes more complicated. The authors have detected a property of the active component of impedance as

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Detection of Concentration Polarization by
Measuring the Electrode Impedance

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a function of $1/\sqrt{\omega}$, which may be used to indicate the presence of concentration polarization. It is assumed that if there is no concentration polarization, any electrode can be simulated in first approximation by one of the circuit diagrams shown in Fig. 1. The impedance of the electrode is given as $Z_o = R_o - jX_o$ (R_o and X_o = active and reactive component, respectively), and the impedance for each diagram of Fig. 1 and the value for $\lim_{\omega \rightarrow 0} R$ are obtained as follows: ✓

$$Z_{o(1)} = \frac{R}{1 + (\omega RC)^2} - j \frac{\omega R^2 C}{1 + (\omega RC)^2} = R_{o(1)} - jX_{o(1)} \quad (1)$$

$$Z_{o(2)} = R - j \frac{1}{\omega C} = R_{o(2)} - jX_{o(2)} \quad (2)$$

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Detection of Concentration Polarization by
Measuring the Electrode Impedance

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$$Z_{0(s)} = \frac{RC_1^2}{\omega^2 R^2 C_1^2 C_2^2 + (C_1 + C_2)^2} - i \frac{\omega^2 R^2 C_1^2 C_2^2 + \omega^{-1} (C_1 + C_2)}{\omega^2 R^2 C_1^2 C_2^2 + (C_1 + C_2)^2} = R_{0(s)} - iX_{0(s)}; \quad (3)$$

$$Z_{0(s)} = \frac{\omega^2 C_1^2 C_2^2 R_1 R_2 (R_1 + R_2) + R_1 C_1^2 + R_2 C_2^2}{\omega^2 (R_1 + R_2)^2 C_1^2 C_2^2 + (C_1 + C_2)^2} - i \frac{\omega C_1 C_2 (R_1^2 C_1 + R_2^2 C_2) + \omega^{-1} (C_1 + C_2)}{\omega^2 (R_1 + R_2)^2 C_1^2 C_2^2 + (C_1 + C_2)^2} = R_{0(s)} - iX_{0(s)}; \quad (4)$$

When an element equivalent to the concentration polarization is introduced into the circuit, the function R_0 becomes infinite. Thus, an unlimited increase of R for $\omega \rightarrow 0$ indicates the presence of concentration polarization, whereas the tendency of R to a limit shows that it is absent. Hence, the dependence of the active component of the electrode impedance R on $1/\omega$ may serve as a criterion for the presence of concentration polarization at the electrode. Professor B. N. Kabanov is thanked for advice. There are 2 figures and 3 references: 2 Soviet and 1 German.

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Detection of Concentration Polarization by
Measuring the Electrode Impedance

S/076/60/034/008/035/039/XX
B015/B063

ASSOCIATION: Akademiya nauk SSSR Institut elektrokhimii
(Academy of Sciences USSR, Institute of Electrochemistry)

SUBMITTED: February 6, 1960

$$\lim_{\omega \rightarrow 0} R_{0(1)} = R;$$

$$\lim_{\omega \rightarrow 0} R_{0(2)} = R;$$

$$\lim_{\omega \rightarrow 0} R_{0(3)} = \frac{RC_1^2}{(C_1 + C_2)^2} < R;$$

$$\lim_{\omega \rightarrow 0} R_{0(4)} = \frac{R_1 C_1^2}{(C_1 + C_2)^2} + \frac{R_2 C_2^2}{(C_1 + C_2)^2} < R_1 + R_2.$$

(5)

(6)

(7)

(8)

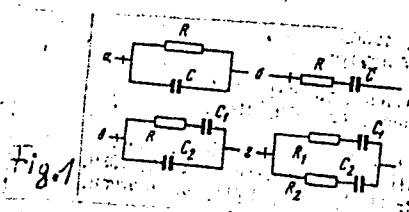


Fig. 1

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S/020/60/135/005/036/043
B004/B075

26.2570

AUTHORS:

Ukshe, Ye. A., Bukun, N. G., and Leykis, D. I.

TITLE:

The Capacity of an Electric Double Layer in Molten Salts

PERIODICAL:

Doklady Akademii nauk SSSR, 1960, Vol. 135, No. 5,
pp. 1183-1186

TEXT: At present, the authors are systematically investigating the capacity of an electric double layer in molten salts. The present paper reports on results so far obtained. The experiments were carried out in an argon atmosphere at 700-800°C. The analyzer consisted of quartz with a polished molybdenum glass cover. Spectroscopically pure graphite served as a polarization electrode, Pb/KCl, NbCl (1:1), 2.5 mole% PbCl₂ as a reference electrode. A melt of KCl and NaCl was used as an electrolyte. The authors investigated the capacity C of a lead- and a cadmium electrode as a function of the potential ψ at 20-200 kc/sec. This function showed a marked minimum, the potential in the case of a capacity minimum depending only on the metal type, but not on frequency and (up to 100°C) not on temperature. Capacity and resistance were dependent on frequency. Here a greater spread of values

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The Capacity of an Electric Double Layer in
Molten Salts

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B004/B075

was found than in solutions, and continued to be so even at high temperatures. Therefore, the data measured at 200 kc/sec were taken as standards. The potential corresponding to the capacity minimum, almost fully coincides with the maximum potential of the electrocapillary curve (Refs. 4, 5). By this means, a determination of the zero potential of metals by measuring the capacity in melts is rendered possible. In order to prove the measured capacities as actually being those of the electric double layer, it is stated that the electrocapillary curve was obtained by double integration of the curve $C = f(\psi)$. Therefore, the following conclusions are drawn: The capacity of the electric double layer in molten electrolytes has a minimum at the zero potential of the metal. The minimum of the function $C = f(\psi)$ corresponds to a capacity of 36-38 microfarads/cm². Capacity increase of the electric double layer with increasing charge of the metal surface can be explained by a deformation of the electron shells of the ions and by a structural densification of the melt (displacement of the holes from the layer near the electrode). With a considerable shift of the potential, capacity appears, caused by a shift of the potential.

LEYKIS, D.I.

Determination of the zero charge potential of silver. Dokl. AN SSSR
135 no.6:1429-1431 D '60. (MIRA 13:12)

1. Institut elektrokhemii Akademii nauk SSSR. Predstavleno akademikom
A.N. Frankinyam.

(Silver)

LEYKIS, D.I.

Zur frage der Impedanzmessungen an festen elektroden."

Report Submitted to the Intl. Committee for Electrochemical Thermodynamics
and Kinetics Rome, Italy 24-29 Sep 1962

UKSHE, Ye.A.; BUKUN, N.G.; LEYKIS, D.I.

Double electrical layer in fused salts. Zhur. fiz. khim. 36
no.11:2322-2328 N'62. (MIRA 17:5)

1. Institut elektrokhimii AN SSSR i Bereznikovskiy filial
Vsesoyuznogo alyuminiyevo-magniyevogo instituta.

S/020/62/142/001/018/021
B145/B101

AUTHORS: Vidovich, G. L., Leykis, D. I., and Kabanov, B. N.

TITLE: Mechanism of anodic oxidation of silver in alkali

PERIODICAL: Akademiya nauk SSSR. Doklady, v. 142, no. 1, 1962, 109-112

TEXT: A previous paper by the authors (DAN v. 124, no. 4, 1959, 855) showed that the anode current in anodic oxidation of silver in alkali at constant potential (0.32 v) decreased in the first minutes after switching on, remained constant for some time, and subsequently continued decreasing. The reasons for this course of the current-time curve are studied in this paper. The overvoltage curve of the $\text{Ag} \rightarrow \text{Ag}_2\text{O}$ process shows a sharp increase of the overvoltage potential at approximately 1.10^4 a/cm^2 , indicating that a diffusion limiting current has been reached. Both the limiting current and its independence of the KOH concentration exclude a concentration polarization as regards the OH^- ion. On the contrary, as was shown by measuring the impedance components of the electrode in alternating current at different frequencies, ω , there is a limitation of diffusion in the arising Ag_2O layer of approximately 5.10^{-7} cm thickness.

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S/020/62/142/001/018/021
B145/B101

Mechanism of anodic oxidation ...

The active impedance component, R_0 , of the Ag electrode increases indefinitely with $1/\omega$. The resistance equivalent to the concentration polarization is approximately 100 times higher than in the concentration polarization of the diffusing ion in solution. R_0 does not depend on the KOH concentration. It may be concluded from these results that diffusion is complicated in the oxide layer. Ag ions diffuse through the Ag_2O layer, and the $Ag \rightarrow Ag_2O$ process takes place on the boundary layer between solution and Ag_2O . This process limited by diffusion does not continue at the same rate until the Ag electrode is completely consumed but passivation occurs, as is shown by the decrease in the anode current. This may also be observed in a shift of the overvoltage curve toward lower $\log i$ values for this period. The measurement of the active component of impedance shows that there exists no concentration polarization for this period. The following is suggested to explain passivation: Accumulation of adsorbed oxygen on the oxide surface layer effects both a decrease of the adsorption rate of oxygen and of the ionization rate of Ag atoms. The resistance of the electrochemical process calculated from the inclination of the overvoltage curve ($\psi - \log i$) and the value i increases

KABANOV, B.N.; LEYKIS, D.I.; KISELEVA, I.G.; ASTAKHOV, I.I.; ALEKSANDROVA,
D.P.

Cathodic introduction of alkali metals into electrodes in aqueous
solutions. Dokl. AN SSSR 144 no.5:1085-1088 Je '62.
(MIRA 15:6)

1. Institut elektrokhemii AN SSSR. Predstavleno akademikom
A.N.Frumkinym.
(Intermetallic compounds) (Electrochemistry)

IEYKIS, D.I.; SEVAST'YANOV, E.S.

Measurement of differential capacitance on solid and liquid electrodes.
Dokl. AN SSSR. 144 no.6:1320-1323 Je '62. (MIRA 15:6)

1. Institut elektirokhemii Akademii nauk SSSR. Predstavleno akad.
A.N.Frumkinym.
(Electrodes)

KASHCHEYEV, V.D.; KABANOV, B.N.; LEYKIS, D.I.

Anode activation of iron. Dokl. AN SSSR 147 no.1:143-145
N '62. (MIRA 15:11)

1. Institut elektrokhemii AN SSSR. Predstavleno
akademikom A.N. Frumkinym.

(Iron)
(Electrodes)

S/062/63/000/001/005/025
B101/B186

AUTHORS: Ukshe, Ye. A., Bukun, N. G., and Leykis, D. I.

TITLE: Effect of the nature of the electrolyte on the capacity of the double layer in molten salts

PERIODICAL: Akademiya nauk SSSR. Izvestiya. Otdeleniye khimicheskikh nauk, no. 1, 1963, 31-36

TEXT: The capacity of the double layer of a molten lead electrode was measured at 18-20 kc/s. The electrode was in a capillary tube and covered with molten alkali halide. The reference electrode used was: Pb|10% by weight PbCl₂ + electrolyte tested. The following compounds were studied as electrolytes: NaCl; NaBr; NaI; NaCl + NaF 1:1; NaCl + KCl 1:1; NaCl + NaI; KCl, KBr, KI, KCl + KF; CsCl, CsI; LiCl; LiCl + 20 mole% LiI. The temperature was 800°C, with NaCl 820°C. Results: The nature of the electrolyte had a marked effect on the structure of the double layer. The alkali halides can be subdivided into two groups. Na and Li halides give high capacity, this being strongly dependent on the nature of the anion: C_{min} ($\mu F/cm^2$) is for NaCl 43, NaBr 52, NaI 75, LiCl 35, Card 1/3

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Effect of the nature of the ...

S/062/63/000/001/005/025
B101/B186

ASSOCIATION:

Institut elektrokhemii Akademii nauk SSSR (Institute of
Electrochemistry of the Academy of Sciences USSR);
Bereznikovskiy filial Vsesoyuznogo alyuminiyevo-
magniyevogo instituta (Berezniki Branch of the All-Union
Institute of Aluminum and Magnesium)

SUBMITTED:

May 9, 1962

Card 3/3

SEVAST'YANOV, E.S.; LEYKIS, D.I.

Adsorption of hexyl alcohol on solid metals. Izv. AN SSSR. Ser.khim.
no.3:450-453 Mr '64. (MIRA 17:4)

1. Institut elektrokhemii AN SSSR.

LEYKIS, D.I.; SEVAST'YANOV, E.S.; KNOTS, L.L.

Change in the impedance components of an electrode with
change of alternating current frequency. Zhur. fiz. khim.
38 no.7:1833-1837 J1 '64. (MIRA 18:3)

1. Institut elektrokhemii AN SSSR.

SEVAST'YANOV, E.S.; LEYKIS, D.I.

Differential capacity measurement on liquid and solid mercury
in methanol. Elektrokhimiya 1 no.2:239-241 F '65.

(MIRA 18:6)

1. Institut elektrokhemii AN SSSR.

ALEKSANDROVA, D.P.; LEYKIS, D.I.

Electric double layer capacitance in concentrated salt solutions.
Elektrokhimiia 1 no.2:241-243 F '65. (MIRA 13:6)

1. Institut elektrokhemii AN SSSR.

LEYKIS, D.I.

Interpretation of the results of impedance measurement in
electrochemical studies. Elektrokhiimiia 1 no.4:472-476 Ap '65.
(MIRA 18:6)

1. Institut elektrokhimii AN SSSR.

LEYKIS, D.I.; VIDOVICH, G.L.

Interpretation of the results of impedance measurement in the
course of anodic oxidation of silver in alkali. Elektrokimiia
1 no.4:477-478 Ap '65. (MIRA 18:6)

1. Institut elektrokhimii AN SSSR.

LEYKIS, K. I. and SEVOSTYANOV, E. S. (Institute of electrochemistry of Academy of Sciences
of USSR)

"Investigation of adsorption of hexyl alcohol on a group of metals"

Report presented at the Intervuz Conference on Electrodeposition of Nonferrous Metals,
Ural Polytechnical Institute im S. M. Kirov, Sverdlovsk, held from 27-30 May 1963.

(Reported in Tsvetnyye Metally, No. 10, 1963, pp. 82-84)
JPRS 24,651 19 May 1964

Leyko, B. P.

AID P - 4030

Subject : USSR/Power

Card 1/1 Pub. 26 - 19/31

Author : Leyko, B. P., Eng.

Title : Performance of boilers without water regulators.

Periodical : Elek. sta., 11, 49-52, N 1955

Abstract : The article describes in detail the defects in the operation of 10 boilers equipped with foreign-made water regulators and their subsequent reconstruction. A special regulator was designed and installed. Its operation is explained. Three diagrams.

Institution : None

Submitted : No date

1st and 2nd GROUPTS

PROCESSES AND PROPERTIES

112

ca

LEYKO, E.

The pharmacological action of nymphalin. III. Action on cardiac tonus and coronary circulation. *Bull. Leyko and Jerzy Modrakowski. Bull. intern. acad. Polon. sci., Cl. med.* No. 5/7, 327-332 (1934); *Rev. ges. Physiol. expl. Pharmacol.* 86, 342; *cl. C. A.* 31, 3571. The glucoside nymphalin from *Nymphes alba* and *Nymphes luteum* was tested on a Starling heart-lung prep. Concn. up to 0.1 mg. per l. of blood produced prolonged increase in tonus and up to 50% increases in coronary flow. The optimum action developed in 20 to 30 min. Concn. between 0.1 mg. and 0.2 mg. per l. of blood were unsafe, and stronger solus. were toxic. J. C. M.

ASB 31A METALLURGICAL LITERATURE CLASSIFICATION

LEYKO, Emil; LIPINSKA, Hanna; SAWICKI, Zdzislaw.

Severe efenal (Luminal) poisoning treated with picrotoxin.
Polski tygod.lek. 10 no.46:1503-1505 14 Nov. '55.

1. Z III Kliniki Chorob Wewn. A.M. L.; kierownik: prof.dr med.
Waclaw Markert i z Zakladu Farmakologii A.M. L. kierownik:
prof. dr med. Emil Leyko. Lodz, ul. Narutowicza 60.

(BARBITURATES, poisoning,
phenobarbital, ther.,picrotoxin)
(POISONING,
phenobarbital, ther.picrotoxin)
(PICROTOXIN, therapeutic use,
phenobarbital, pois.)

P.T.A. LEYKO, Ya.

Meet. & Elec. Engineering

535

629 136

Leyko J. The Calculation of the Framework Fuselage of a Trapezoidal Cross Section.

„Obliczenie kadłuba kratowego o trapezowym przekroju poprzecznym” Technika Lotnicza, No. 2, 1950, pp. 60—66, 17 figs

Calculation of loading of the longitudinal plane frameworks. Calculation of the force in the diagonal member of the zero partition. Calculation of forces in members. The case when one of the members of the longitudinal framework has been replaced by a diagonal member of the inside partition. The article gives an original development and a practical application of the orthodox method.

2868

539.38 : 669-462 : 517.2

Leyko J. Tensions in a Rotating Cylindrical Shell Loaded on the Edge with Longitudinal Forces.

Napięcia w powłoce obrotowej cylindrycznej na brzo-
wach wzdłużnymi" (Prace Inst. Mechan. No. 30, Warszawa, 1952, 1957,
16 pp., 16 figs., 2 tabs.

The author discusses the bending of a beam in the form of a cy-
lindrical shell. The case analysed deals with reactions distributed along
small sections of the circumference. A solution is advanced which, on
the basis of the general theory of cylindrical shells, fulfills all edge
conditions. A numerical example is given and a comparison made be-
tween the tensions calculated and the corresponding tensions arising
from the elementary theory of bending.

LEYKO, J.

Problem of dynamic coupling of the torsion vibrations with the bending vibrations of a fast rotating crankshaft. p. 273.
ARCHIWUM BUDOWY MASZYN, Warszawa, Vol. 1, no. 3, 1954.

SO: Monthly List of East European Accessions, (SEAL), LC, Vol. 4, no. 10, Oct. 1955,
Uncl.

LEYKO, J.

6

CH ✓ The role of vanadyl sulfate in the process of catalytic oxidation of sulfur dioxide. S. Weychert and J. Leyko. (Inst. Tech., Warsaw). Bull. acad. polon. sci. Classe III, 3, 329-332 (1955) (in English).—The synthesis and the decomposition of VOSO_4 were studied between 300 and 630°, by means of a thermobalance. V_2O_5 was treated with air contg. 10-45% SO_2 with and without the addn. of SO_3 . Not all of the V_2O_5 present could be converted to VOSO_4 ; the degree of conversion rose continually with temp. VOSO_4 began to decomp. at about 565°; decompn. was complete at about 580°. The initial rate of formation of SO_3 was of the same order as the rate of oxidation of SO_2 . After a certain degree of conversion the rate decreased, and finally decompn. practically ceased. By the action of a mixt. of SO_2 , SO_3 , and O_2 the whole catalyst mass can be converted to V_2O_5 , provided that the equil. conditions for VOSO_4 stability are greatly exceeded. When V_2O_5 is reduced by pure SO_2 at low temps., VOSO_4 is first formed; later this changes to V_2O_5 . At higher temps. V_2O_5 is formed directly. K. E. Hayes

[Handwritten signature]

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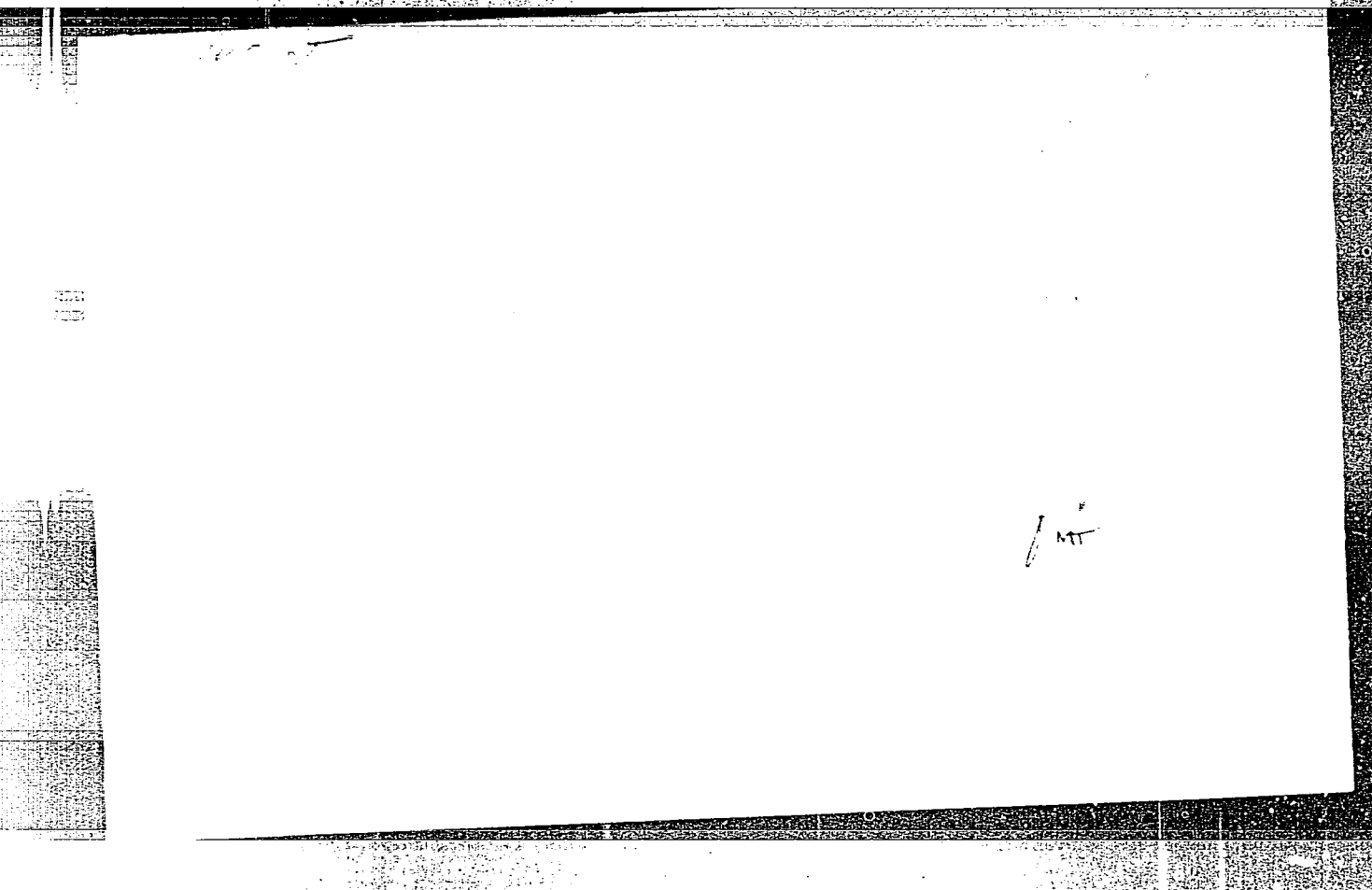
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APPROVED FOR RELEASE: Monday, July 31, 2000

CIA-RDP86-00513R000929720C

1.1 440.5

DATA,

Stefan Weychert and Jadwiga Leyko: "Synthesis and Decomposition of Vanadyl Sulfate in the Process of Contact Oxidation of Sulfur Dioxide, Roczniki Chemii, Vol 30, No 1, Warsaw, 1956. Published from the Research Laboratory of the Technology of Inorganic Chemistry, Warsaw Polytechnic, 24 Jan 55.

LEYKO, Jerzy (Lodz)

Stability of an orthotropic cone sector shaped shell compressed
along the generatrices. Archiw bud masz 3 no.4:445-460 '61.

BRETSZNAJDER, Stanislaw; LEYKO, Jadwiga; BLUM, Aleksander

On the rate of nuclei formation of a new solid phase in the reaction of thermal dissociation of magnesium carbonate. Roczniki chemii 35 no.5: 1477-1486 '61.

1. Department of Technological designing, Technical University, Warsaw and Department of Physicochemical Problems of Technology, Institute of Physical Chemistry, Polish Academy of Sciences, Warsaw.

LEYKO, Jerzy; MLOTKOWSKI, Andrzej (Lodz)

Bending of an axially symmetric, orthotropic, circular plate of variable rigidity, loaded by a couple of forces applied to the center. Archiw bud masz 9 no.3:375-391 '62.

LEYKO, Jerzy

The state of stress in the case of shrink fit of two cylindrical shells reinforced by longitudinal ribs. Archiw mech
15 no.4:547-564 '63

1. Technical University, Lodz

IEYKO, Jerzy (Lodz)

Approximate differential equations of shrink joined cylindrical shells. Archiw bud masz 11 no. 1:9-13 '64.

LEYKO, J.; MACIEJEWSKI, M.

Nucleation rate in the thermal dissociation of silver carbonate.
Bul chim PAN 12 no.4:273-274 '64.

1. Department of Basic Physicochemical Problems in Technology,
Institute of Physical Chemistry, Polish Academy of Sciences,
Warsaw. Presented by S. Bretsznajder.

LEYKO, J.

Equilibrium studies on the $\text{H}_2\text{S}-\text{NH}_3-\text{H}_2\text{O}$ system. Pt. 1. *Bul chim PAN* 12 no.4:275-276 '64.

1. Department of Basic Physicochemical Problems in Technology,
Institute of Physical Chemistry, Polish Academy of Sciences,
Warsaw. Presented by S. Bretsznajder.

LEYKO, J.

Dissolution of sulfur in aqueous ammonium sulfide solutions.
Bul chim PAN 12 no.6:441-443 '64.

1. Department of Basic Physico-Chemical Problems in Technology
of the Institute of Physical Chemistry of the Polish Academy of
Sciences, Warsaw. Submitted February 25, 1964.

LEYKO, J.; PIATKIEWICZ, J.

Equilibrium studies on the $\text{H}_2\text{N}-\text{NH}_2-\text{H}_2\text{O}$ system. Pt.2.
Bul chim PAN 12 no.6:445-446 '64.

1. Department of Basic Physico Chemical Problems in Technology of the Institute of Physical Chemistry of the Polish Academy of Sciences, Warsaw, and Department of Technological Design of Warsaw Technical University, Submitted February 25, 1964.

24.4200

P/032/60/007/002/002/003
D259/D304

AUTHOR:

Leyko, Jerzy (Łóź)

TITLE:

Non-linear problem of the equilibrium of the shell in the form of a cylindrical segment subjected to shear and normal load

PERIODICAL: Archiwum budowy maszyn, v.7, no.2, 1960, 199-211

TEXT: This is an approximate solution of a non-linear problem of the equilibrium of a thin elastic shell in the form of a segment of a circular cylinder freely supported on the edges and subjected to shear and normal load, uniformly distributed on the whole surface. It is solved by means of differential equations for shells of small curvature and the relation is established between the mean shearing forces on the edges of the shell and the deflection of the middle surface. There are 5 figures and 2 Soviet-bloc references.

Card 1/2

LEYKO, Jerzy, prof, dr

"Rheological properties of model materials" by [doc. dr inz.]
J.T. Pindera. Reviewed by Jerzy Leyko. Przegl mech 22 no.9:
291 10 My '63.

MIROSHNICHENKO, Ye.A.; LEYKO, V.P.; LEBEDEV, Yu.A. (Moscow)

Semimicrocalorimeter. Zhur. fiz. khim. 38 no.4:1054-1055 Ap '64.
(MIRA 17:6)

1. Akademiya nauk SSSR, Institut khimicheskoy fiziki.

LEYKO, V. S., Candidate Tech Sci (diss) -- "Problems of using a turbo-clutch for electrical reversing construction winches". Moscow, 1959. 12 pp (Min Higher Educ USSR, Moscow Order of Labor Red Banner Construction Engineering Inst im V. V. Kuybyshev), 200 copies (KL, No 25, 1959, 134)

LEYKO, WANDA

2

Oxidation-reduction potential of the cysteine-cystine system. Wanda Leyko. Bull. soc. chim. et lectes Lids Classe III, 3, No. 14, 14 pp. (1932) (in English).—The work (Chosh, et al., C.A. 26, 3976) on the potential of partially (electrolytically) reduced cystine (RSSR, RSH) in the cell: Hg|cystine, cysteine, HCl||saturated KCl||agar||calomel half cell, was repeated. O was removed by a current of N washed 3 times with alk. pyrogallol. In the equation $E = E_0 - 0.059 \text{ pH} - 0.0295 \log \frac{[\text{RSH}]^2}{[\text{RSSR}]}$, E_0 is $+0.091 \text{ v.}$ in satisfactory agreement with G.'s values of $+0.079$ to $+0.085$. By using a Cu electrode E_0 is -0.021 v. The difference is significant. The suggestion of Barron, et al. (C.I. 23, 2729) that the potential observed is due to reaction of cysteine with metal ions was examd. Cu was detd. in the electrode soln. by dithizone and cysteine by the method of Lavine (C.A. 29, 4393). The equation $E = E_0 - 0.059 \text{ pH} - 0.0295 \log \frac{[\text{RSH}]^2}{[\text{RSSR}][\text{Cu}]} (I)$ is obeyed, and the value of E_0 is approx. $+0.03 \text{ v.}$ To obtain an independent check on the assumptions of I, 0.007 M cysteine in 0.1 N HCl was titrated with CuCl₂ potentiometrically (the metal used in the electrode wire is not specified). E values did not fit I unless the concn. of (RS)₂Cu (II) was multiplied by a factor which decreased linearly from 1 to 0.3 as II increased from zero. E_0 then = $+0.074 \text{ v.}$, in satisfactory agreement with the value 0.06 v. above. Reversible complexes of cysteine with Cu or other metals at concns. of the order of 10^{-4} M are responsible for the observed oxidation-reduction potentials of cysteine both *in vitro* and *in vivo* and the concn. of cysteine does not have an influence on these potentials. R. B. D.

BI

FILIPOWICZ, B.; LEYKO, W.; WIECKOWSKI, W.

Nucleic acids of the pancreas. Acta physiol. polon. 3 Suppl. 3: 261-264 1952. (CML 24:1)

1. Of the Institute of General and Physiological Chemistry (Head--Prof. B. Filipowicz, M.D.) of Lodz Medical Academy.

FILIPOWICZ, B.; LEYKO, W.

Determination of adenine in tissues by polarographic method. Acta physiol. polon. 3 Suppl. 3: 264-266 1952. (CLML 24:1)

1. Of the Institute of General and Physiological Chemistry (Head--Prof. B. Filipowicz, M.D.) of Lodz Medical Academy.

Foreign relations of the United States in 1965
Foreign relations of the United States in 1965
Foreign relations of the United States in 1965

LEWKO, W.; PANUSZ, H.

Polarographic method of determination of adenine compounds in the blood. Polski tygod. lek. 9 no.34:1057-1063 23 Aug 54.

1. Z Zakladu Interny Polowej Akademii Medycznej w Lodzi; kierownik:
prof. dr med. A.Himmel.

(BLOOD,
adenine cpds., polarography)
(ADENINE, in blood,
determ., polarography)
(POLAROGRAPHY,
of adenine in blood)

LEYKO, W.; GROSS, S.

Adenine compounds in human blood. I. Polarographic and spectrophotometric determination of adenine in deproteinized blood.
Acta biochim.polon. 2no:2:155-168 1955.

1. Z Zakladu Biochemii Uniwersytetu Lodzkiego: Kierownik: prof.
dr A. Dmochowski i instytutu medycyny Pracy w Lodzi, Dyrektor:
prof. dr E. Paluch.

(ADENINE, in blood,
determ., polarography & spectrophotometry in
deproteinized blood)

(BLOOD,
adenine, polarography & spectrophotometry in
deproteinized blood)

(POLAROGRAPHY,
of blood adenine, after deproteinization)

Leyko, W.

2820. Nephelometric determination of adenine and guanine. W. Leyko and B. Filipowicz (Akad. Med., Lodz, Poland). *Roczn. Chem.*, 1955, 29 (4), 1095-1101.—The adenine content of a pancreas hydrolysate is determined polarographically, and an aliquot containing 24 μ g of adenine is used for the nephelometric determination of total adenine and guanine according to the method of Groves and Kober (*J. Amer. Chem. Soc.*, 1915, 37, 3430). The guanine content of the aliquot is read from an empirical curve relating turbidity with guanine content of a series of solutions containing 24 μ g of adenine and 24 to 83 μ g of guanine. R. Tauscor

WANDA, LEYKO

E-3

Poland/ Organic/Chemistry - Naturally occurring substances
and their synthetic analogs

Abs Jour : Referat Zhur - Khimiya, No 4, 1957, 11882

Author : Leyko Wanda, Filipowicz Bronislaw
Title : Spatial Structure of Nucleic Acids

Orig Pub : Przestrzenna budowa kwasow nukleinowych. Postepy biochem., 1956,
2, No 1, 61-75 (Polish)

Abstract : A review. Bibliography 45 references.

Card 1/1

LEYKO, WANDA

Polarographic Method of the determination of adenine in tissues.
Pt. 2 - Deamination rate of adenine nucleotides in human pancreas, 1957, 6 p.
Pt. 3 - The ratio of guanine to adenine in nucleic acids of human pancreas, 1957, 7 p.
(Lodzkie Towarzystwo Naukowe. Wudzial III: Nauk Matematyczno-Przyrodniczych.
Bulletin, v. 8, ((nr)) 5-6), Lodz, Poland)

SO: Monthly List of East European Accessions (EEAL) LC, Vol. 6, No.9, Sept, 1957 Uncl.

LEYKO, WANDA.

Związki adeninowe trzustki ludzkiej. (Wyd. 1) Lodz (Panstwowe Wydawn. Naukowe)
1957. 40 p. (Lodzkie Towarzystwo Naukowe. Wydział III., Prace, nr. 46)
(Adenine compounds of the human pancreas. 1st ed. English and Russian Summaries
bibl., tables)

SO: Monthly Index of East European Accession (EEAI) LC Vol. 7, No. 5, 1958

LEYKO, W.

B-12

POLAND/Physical Chemistry - Electrochemistry.

Abs Jour: Referat. Zhurnal Khimiya, No 3, 1958, 7325.

Author : W. Leyko, E. Sempinska.

Inst : Academy of Sciences of Poland.

Title : Polarographic Behavior of Kinetin.

Orig Pub: Bull. Acad. polon. sci., 1957, Cl. 2, 5, No 3, 75-78.

Abstract: Kinetin (6-furfurylmethylaminopurine) (I) separated from DNC
[DNC is in the original. Translator does not know what it is.]
and prepared synthetically produces a wave on the background of
0.1 n. HCl with E_1 close to E_2 of adenine. The initial I concen-
tration was computed by adding the standard adenine solution.
The results of the polarographic determinations are compared with
the results of spectrophotometrical measurements and they agree
well.

-22-

Card : 1/1

LEYKO, W.

SCIENCE

Periodical: KOSMOS. SERIA A: BIOLOGIA. Vol. 8, no. 5, 1957. In English.

LEYKO, W. Polarographic method of the determination of adenine in tissue.
Pt. 2. The deamination rate of adenine nucleotides in human pan-
creas. p. 1.

Monthly List of East European Accessions (EAI), IC, Vol. 8, No. 3, May 1959
Unclass.

LEYKO, W.

SCIENCE

Periodical: KOSMOS. SERIA A: BIOLOGIA. Vol. 8, no. 6, 1957. In English.

LEYKO, W. Polarographic method of the determination of adenine in tissue.
Pt. 3. The ratio of guanine to adenine in nucleic acids of human
pancreas. p. 1.

Monthly List of East European Acessions (EEAI), LC, Vol. 8, No. 3, May 1959
Unclass.

LETKO, Wanda; GROSS, Maria; FILIPOWICZ, Bronislaw

Adenine compounds in human blood; comparison of adenine levels in arterial and oxygenated venous blood. Polskie arch. med. wewn. 39 no.1:13-18 1959.

1. Z Zakladu Chemii Fizjologicznej A.M. w Lodzi Kierownik: prof. dr B. Filipowicz. Adres autora: Lodz, Narutowicza 68, Kat. i Zakl. Chemii Fizjologicznej

(ADENINE, in blood, comparison in arterial & oxygenated venous blood (Pol))

LEYKO, Wanda, Dr.; BUCZEK, Lidia, mgr biologii

Structure of the hemoglobin. Postepy biochemii 6 no.4:395-407 '60.
(HEAI 10:3)

1. Adiunkt Zakladu Chemii Fizjologicznej Akademii Medycznej w Lodzi
(for Leyko)
(HEMOGLOBIN)

FILIPOWICZ, Bronislaw; LORENC, Jozefa; LEYKO, Wanda

Comparison of adenine levels in arterial and venous blood. Polskie
arch.med.wewn. 30 no.3:373-377 '60.

1. Z Zakladu Chemii Fizjologicznej A.M. w Lodzi. Kierownik: prof.
dr B. Filipowicz.

(ADENINE blood)

LEYKO, Wanda

Rejuvenation of erythrocytes. Postepy biochem 7 no.3:357-370 '61.

(ERYTHROCYTES) (BLOOD BANKS)

LEYKO, Wanda

The primary structure of proteins. Postepy biochem. 8 no.3:305-315
'62.

(PROTEINS chem)

LEYKO, Wanda

Modern views on the erythrocytes metabolism. Postepy biochem. 8
no.4:511-522 '62. (ERYTHROCYTES)

LEYKO, Wanda; DMOCHOWSKI, Antoni; BUCZEK, Lidia; GONDKO, Roman

Obtaining pure hemoglobin from human blood. Nauki matem
przyrod Lodz no.12:37-49 '62.

1. Katedra Biochemii, Uniwersytet, Lodz.

*

LEYKO, W.; DMOCHOWSKI, A.; BOLANOWSKA, W.; LORENC, J.

The effects of nitrite on adenine compounds in the human blood hemolysates.
Postepy biochem. 8 no.4:554-555 '62.

1. Z Zakladu Biochemii UL i Zakladu Chemii Fizjologicznej AM w Lodzi.
(NITRITES) (ADENINE) (HEMOLYSIS)

KIDIN, I.N.; LEYKOVSKIY, K.K.; PIGUZOV, Yu.V.; FADINA, L.V.

Investigating the isothermal decomposition of austenite by the
internal friction method. Fiz. met. i metalloved. 18 no.2:16-
317 Ag '64. (MIRA 18:8)

1. Moskovskiy institut stali i splavov.

LEYKOVSKIY, M.M. , okulist

Active detection of glaucoma. Zdrav. Belor. 6 no. 10:57 0 '60.
(MIRA 13:10)

1. Iz Volozhinskoy rayonnoy bol'nitsy (glavnyy vrach S.Z. Kipel'),
nauchnyy rukovoditel' raboty - prof. M.M. Zolotareva.
(VOLOZHIN DISTRICT—GLAUCOMA)

LEYKOVSKIY, M.M.

Detection of glaucoma in the White Russian Republic. Vest.oft.
no.4:4-6 '61. (MIRA 14:11)

1. Volozhinskaya rayonnaya bol'nitsa (nauchnyy rukovoditel' -
zav. kafedroy glaznykh bolezney Belorusskogo instituta uso-
vershenstvovaniya vrachey prof. M.M. Zolotareva).
(WHITE RUSSIA--GLAUCOMA)

LEYKOVSKIY, M.M.

Adrenaline test in the diagnosis of glaucma. Zdrav. Bel. 9
no.8476-77 Ag'63 (MIRA 17:3)

1. Iz Volozhinskoy rayonnoy bol'nitsy (glavnyy vrach S.Z. Kipel'); nauchnyy rukovoditel' - zaveduyushchiy kafedroy glaznykh bolezney Belorusskogo instituta usovershenstvovaniya vrachey - prof. M.M. Zolotareva.

RUDBERG, I.M.; LEYKOVSKIY, M.M.

Recognition of shellac calculi of the stomach. Vest. rent.
1 rad. 38 no.1:67 Ja-F'63. (MIRA 16:10)

1. Iz Volzhinskoy rayonnoy bol'nitsy (glavnyy vrach S.Z.
Kipel')

*

S/058/62/000/005/003/119
A160/A101

AUTHORS: Leykum, V. I., Oleynik, B. N., Sozinov, G. Ya.
TITLE: The measurement of the heat conductivity of semiconductors
PERIODICAL: Referativnyy zhurnal, Fizika, no. 5, 1962, 12, abstract 5A131
("Tr. in-tov Kom-ta standartov, mer i izmerit. priborov pri Sov.
Min. SSSR", 1961, no. 51 (III), 131-137)

TEXT: An analysis of the methods of measuring the heat conductivity of semiconductors leads to the conclusion that the best developed methods are the stationary-flow-type ones. A description is given of an installation designed for measuring the heat conductivity and based on the method of "two plates". A flat heater is placed between the two test pieces which are 20 mm long and 10 mm in diameter, and junctions of differential thermocouples are mounted in the pieces. The whole system is placed in an updraft furnace. Two measuring units were designed, one for temperatures from -180 to +20°C, and the other for temperatures from 20 to 800°C. The temperature in the units is set for 4 - 6 hours. The installation was tested within a temperature range of -180 to +500°C. The

Card 1/2

40504

S/263/62/000/013/008/015
1007/1207

21.2532
AUTHORS: Leykum, V. I., Oleynik, B. N., Sozinov, G. Ya.

TITLE: Measurement of heat conductivity of semiconductors

PERIODICAL: Referativnyy zhurnal, otdel'nyy vypusk. 32. Izmeritel'naya tekhnika, no. 13, 1962, 46, abstract 32.13.335. (Tr. in-tov Kom-ta standartov, mer i izmerit. priborov pri Sov. Min. SSSR, no. 51 (111), 1961, 131-137)

TEXT: This is a brief review of various methods for measuring heat conductivity. The method of steady heat flow was used for measuring the heat conductivity of semiconductors. There is a description of apparatus, based on this method and developed at the VNIIM, for measuring the heat conductivity of two semi-conductor test specimens, 10 mm in diameter and 200 mm long. The measuring device consists of four separate units: a vacuum furnace for the 20 to 800°C range, a low-temperature unit for the -80 to +20°C range, a vacuum-generating unit and a measuring unit. The apparatus, based on the "two plate" method, works in the following manner: the test specimen is cut into two approximately equal parts between which the flat heating element is placed. Each of the two parts is provided with a differential thermocouple for measuring the temperature gradient along their height. The temperature level at which measurements may be carried out is determined according to the readings

Card 1/2

AUTHORS:

TITLE:

SOURCE:

S/589/62/000/063/011/021
E194/E436
Ipatov, Yu.S., Leykum, V.I., Oleynik, B.N.,
Patovskaya, Z.K.

Instruments for measuring thermal conductivity
USSR. Komitet standartov, mer i izmeritel'nykh
priborov. Trudy institutov Komiteta. no.63(123).
Moscow, 1962. Issledovaniya v oblasti teplovykh i
temperaturnykh izmereniy. 143-150

TEXT: In 1960 a need was pointed out for instruments to measure thermal conductivity quickly, though not necessarily with great accuracy. For this purpose the here described instrument A-21 is suitable. The cylindrical specimen is in contact with an insulated heater. The measurements are carried out with a constant current passing through specimens is a function of their thermal conductivity, providing they have the same length. In the base of the instrument is a transformer which supplies the heater. Microammeter M-95 measures the output from the thermocouples in the contact plates at the top and bottom of the test

Card 1/3

PROVED FOR

CIA-RDP86-00513R0009297

Instruments for measuring thermal ... S/589/62/000/063/011/021
E194/E436

piece. The influence on the readings of the pressure on the specimen, the air temperature around the specimen and the cooling temperature were studied, using test pieces 15 mm diameter and 20 mm long made of fused quartz or of fluoroplast-4 (teflon). These specimens are longer than the optimum length of 5 to 10 mm and so the maximum stabilisation time could readily be determined, with quartz it was up to 125 minutes. Similar results were obtained with fluoroplast-4. The specimen diameter is not critical but it will probably be convenient to use a value of 10 to 15 mm. The mean standard error was $\pm 1\%$ and the overall error of measuring the coefficient of thermal conductivity of a specimen 10 to 15 mm diameter and 5 to 10 mm long is about $\pm 2\%$ for a large number of tests or about $\pm 3\%$ for two or three measurements. These values apply for thermal conductivities in the range 0 to 1 W/(metre degree) but for materials of greater thermal conductivity the errors will be greater and indeed the A-21 instrument is not recommended for them. A somewhat improved version, type A-22, is of very similar performance. In order to extend the range of thermal conductivity that can be measured,

Card 2/3

L 15177-66 EWT(m) DIAAP

ACC NR: AP6001143 SOURCE CODE: UR/0367/65/002/003/0402/0408

AUTHOR: Baskova, K. A. ; Vasil'yev, S. S. ; Rudenko, N. P. ; Sevast'yanov, A. I. ; Khamo-
Leyla, M. A. ; Shavtvalov, L. Ya. 4/2 B

ORG: Institute of Nuclear Physics, Moscow State University (Institut yadernoy fiziki
Moskovskogo gosudarstvennogo universiteta)

TITLE: Investigation of the ^{19, 44, 55}radiation of ¹¹⁷Cd

SOURCE: Yadernaya fizika, v. 2, no. 3, 1965, 402-408

TOPIC TAGS: cadmium, beta spectrum, half life, isotope separation, indium

ABSTRACT: ¹¹⁷Cd was obtained from the reaction ¹¹⁶Cd(d, p). As a result of the investigations conducted it is shown that the half-life of ¹¹⁷Cd is about three hours. The half-life of 50 min previously ascribed erroneously to ¹¹⁷Cd is, apparently, that of ¹¹⁶In obtained from the reaction ¹¹⁶Cd(d, 2n). The beta-spectrum of ¹¹⁷Cd (3 hr) was investigated on a beta-spectrometer with a magnetic lens. The upper boundaries of the partial beta-spectra have the energy of 670; 1290; 1800; and 2200 kev. The value of log ft proved to be equal to 4.9; 6.7; 6.9; and 7.6, respectively. The results presented, as well as the investigations of the $\beta\gamma$ -coincidences made it possible to construct a decay scheme of ¹¹⁷Cd which differs substantially from that in the literature. Authors express their gratitude to Yu. A. Vorob'yev, V. S. Zazulin, N. S. Kirnichev, and M. R. Akhmed for assistance in the work. Orig. art. has: 7 figures and 1 table.

Card 1/1 SUB CODE: 20, 18 / SUBM DATE: 19Feb65 / ORIG REF: 001 / OTH REF: 012

KRAKHMAL'NIKOVA, G.A.; KIRENKOV, I.I.; Primali uchastiye: LEYKUM, V.Ye.;
FEDOROV, Ye.V.; UGOL'NIKOV, V.I.; SEMENOVA, L.I.

Spectropyrometric unit designed by the All-Union Research Institute
of Metrology. Izv.tekh. no.5:18-19 My '62. (MIRA 15:6)
(Pyrometers)

IVENSKIY, Yu.N., inzh.; LEYMAN, A.A., inzh.; GOTOVSKIY, A.S., inzh.

Calculation and design of generator-type contactless track
switches. Elektrotehnika 36 no.4:20-23 Ap '65. (MIRA 18:5)

LEYMAN, A.V.; KUZ'MICHEV, A.Ye., kand. tekhn. nauk, nauchnyy red.;
SHIROKOVA, G.M., red. izd-va; NAUMOVA, G.D., tekhn. red.

[Manufacturing and using prestressed concrete elements in the
Sverdlovsk Economic Region] Opyt izgotovleniia i primeneniia
predvaritel'no napriazhennykh zhelezobetonnykh konstruktsii v
Sverdlovskom raione. Moskva, Gos. izd-vo lit-ry po stroit.,
arkhit. i stroit. materialam, 1961. 50 p. (MIRA 15:3)
(Prestressed concrete construction)

KUZNETSOV, A.V., inzh.; LEYMAN, A.V., inzh.

Precast concrete at industrial construction projects in the
Central Urals. Prom. stroi. 41 no.8:24-26 Ag '64. (MIRA 17:11)

1. Glavreduralstroy.

LEYMAN, A.V., inzh.

Preparation of prestressed concrete elements in the Sverdlovsk
Economic Region. Bet. 1 zhel.-bet. no. 3:118-130 Mr '61.
(MIRA 14:5)
(Sverdlovsk Province—Prestressed concrete construction)

LEYMAN, A.V., inzh.

Prefabrication techniques in industrial construction in
Sverdlovsk Province. Prom.stroi. 39 no.8:4-6 '61. (MIRA 14:9)

1. Upravleniye stroitel'stva sovmarkhoza Sverdlovskogo
ekonomicheskogo rayona.
(Sverdlovsk Province—Construction industry)
(Precast concrete construction)

KHARCHEV, A.G., kand.filosof.nauk; LEYMAN, I.I.

Methodological seminars held by Leningrad scientists. Vest.
AN SSSR 34 no. 1:45-47 Ja '64. (MIRA 17:5)

1. Leningradskaya kafedra filosofii AN SSSR.

LEYMAN, M. Zh., VAVILOVA, M. P., SOFIYEV, M. S. and GEORGIYEVSKAYA, N. A.

"The Rational Times for Spring Antirelapse Treatment for Malaria in
Connection With Variations for Incidence in Uzbekistan", Med. Paraz. i Paraz. Bolez.,
Vol. 17, No. 5, pp 412-19, 1948.

LEYMAN, N.P.; STRAKHOVA, L.M.; DENISENKO, E.V. (Saratov)

Prevention of poliomyelitis with nonspecific immunization methods
and their comparative effectiveness. Vop.okh.mat. i det. 1 no.1:
33-36 Ja-F' 56. (MIRA 9:9)

1. Zasluzhennyi vrach RSFSR (for Leyman)
(POLIOMYELITIS--PREVENTIVE INOCULATION)

UTKOP, I.N.; ILYIN, P.I.; GAL'NIKOV, A.A.

Prospect for the natural gas supply for the Latvian S.S.R.
Gen. prom. no. 10:42-11 0 '61. (MIRA 11:11)
(Latvia--Gas, Natural)

AL'TMAN, A.M., inzh.; KONTOROVICH, S.O., inzh.; LEYMAN, P.P., inzh.

Trailer houses for pipeline construction workers. Stroi.truboprov.
3 no.9:20-22 S '58. (MIRA 11:12)

(Labor and laboring classes--Dwellings)
(Automobiles--Trailers)

80V/123-59-16-65532

Translation from: Referativnyy zhurnal. Mashinostroyeniye, 1959, Nr 16, p 266 (USSR)

AUTHORS: Apartsev, A.S., Kann, A.V., Kontorovich, S.O., Leyman, P.P.

TITLE: A New Technology of Constructing Delivery Pipelines of Semi-Metallic Ferro-Concrete Tubes

PERIODICAL: Str-vo truboprovodov, 1958, Nr 11, 13 - 16

ABSTRACT: The economic and operational superiority of employing semi-metallic tubes (T) is stated, consisting of an outer pre-strained reinforced concrete shell and a thin-walled metallic inner sleeve, which increases the service life of the T. The Leningrad "Barrikady" Plant finished the tests with a pilot KZhB-67 machine for the manufacture of pressureless reinforced concrete T of 900 mm in diameter directly in the ditch. The technology of constructing delivery pipelines, worked out by the State Institute for the Designing of Special Enterprises for the Gas Industry "Giprospetsgaz", of semi-metallic T by the method of continuous molding is described. The expediency of applying shells of pre-strained reinforced

Card 1/2

14(5)

SOV/95-59-3-3/14

AUTHORS:

Al'tman, A.M., Kontorovich, S.O., Leyman, P.P., Engineers

TITLE:

Mobile Production Centers on Pipeline Tracks
(Peredvizhnyye proizvodstvennyye bazy na trassakh)

PERIODICAL:

Stroitel'stvo truboprovodov, 1959, Nr 3, pp 5-10 (USSR)

ABSTRACT:

The vast production program of the 7-Year Plan has rendered necessary the establishment of field production centers which should be in a position to provide living accommodations, workshops, and stores for supplies and material required on the construction sites by the working teams. The task of setting up these centers, or service stations, is complicated by the conditions under which they have to function. Located often in remote areas, far from any industrial centers, RR lines or highways, they are intended for temporary stay only, being called upon to move along the track as construction is progressing. Giprospekgaz has worked out a project pertaining to complete typical production centers which answer all the requirements of the pipeline construction and can be erected or dismantled in 10-15 days. The general layout of a production center is shown on diagram 1 as consisting of

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3 main sections: construction area, motor pool for 50 automobiles and stores for fuel and lubricating oils. Table 1 and 2 give a breakdown of the field production center by departments or units, indicating their capacity and kind of constructions. The center is composed of 31 buildings. As a rule all constructions are made for assembling and dismantling on the spot for ready transportation. Typical centers make it possible to organize production from locally available raw material and construction material; they can produce sections of electrotechnical, sanitary, and technological installations and perform maintenance and repair jobs on automobiles and machines; they also act as distributors of fuel, oil, and such materials as they receive for storage. The rest of the article deals with typical buildings assembled from prefabricated parts and panels, made of metal and wood. Dismantled constructions and equipment can be loaded on trucks and trailers, or on RR platforms for transportation. A production center needs about 4,500 m² of open space for putting

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up buildings and 1,000 m² for sheds. It is estimated that it takes 50 men 15 days to put up the principal building complex of a field production center. There are 4 tables, 2 sketches, and 3 diagrams.

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